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(54) Chemical mechanical polishing slurry useful for copper substrates

(57) A chemical mechanical polishing slurry comprising a film forming agent, an oxidizer such as urea hydrogen peroxide, a complexing agent such as ammonium oxalate or tartaric acid, an abrasive, and an op-

tional surfactant. Also disclosed is a method for using the chemical mechanical polishing slurry composition to remove copper alloy, titanium, and titanium nitride containing layers from a substrate.

EP 0 846 742 A2

Description

BACKGROUND OF THE INVENTION(1) Field of the Invention.

This invention concerns a chemical mechanical polishing slurry including a complexing agent, at least one oxidizer, at least one abrasive, and a film forming agent. The chemical mechanical polishing slurry is useful for polishing metal layers and thin-films associated with semiconductor manufacturing. More particularly this invention concerns a chemical mechanical polishing slurry that is especially adapted for polishing multiple metal layers and thin-films where one of the layers or films is comprised of copper or a copper containing alloy.

(2) Description of the Art.

Integrated circuits are made up of millions of active devices formed in or on a silicon substrate. The active devices, which are initially isolated from one another, are interconnected to form functional circuits and components. The devices are interconnected through the use of well-known multilevel interconnections. Interconnection structures normally have a first layer of metallization, an interconnection layer, a second level of metallization, and sometimes a third and subsequent level of metallization. Interlevel dielectrics such as doped and undoped silicon dioxide (SiO_2), are used to electrically isolate the different levels of metallization in a silicon substrate or well. The electrical connections between different interconnection levels are made through the use of metallized vias. U.S. Patent No. 4,789,648, which is incorporated herein by reference, describes a method for preparing multiple metallized layers and metallized vias in insulator films. In a similar manner, metal contacts are used to form electrical connections between interconnection levels and devices formed in a well. The metal vias and contacts may be filled with various metals and alloys including titanium (Ti), titanium nitride (TiN), tantalum (Ta), aluminum copper (Al-Cu), aluminum silicon (Al-Si), copper (Cu), tungsten (W), and combinations thereof. The metal vias and contacts generally employ an adhesion layer such as titanium nitride (TiN) and/or titanium (Ti) to adhere the metal layer to the SiO_2 substrate. At the contact level, the adhesion layer acts as a diffusion barrier to prevent the filled metal and SiO_2 from reacting.

In one semiconductor manufacturing process, metallized vias or contacts are formed by a blanket metal deposition followed by a chemical mechanical polish (CMP) step. In a typical process, via holes are etched through an interlevel dielectric (ILD) to interconnection lines or to a semiconductor substrate. Next, a thin adhesion layer such as titanium nitride and/or titanium is generally formed over the ILD and is directed into the etched via hole. Then, a metal film is blanket deposited over the adhesion layer and into the via hole. Deposition is continued until the via hole is filled with the blanket deposited metal. Finally, the excess metal is removed by chemical mechanical polishing, (CMP) to form metal vias. Processes for manufacturing and/or CMP of vias are disclosed in U.S. Patent Nos. 4,671,851, 4,910,155 and 4,944,835.

In a typical chemical mechanical polishing process, the substrate is placed in direct contact with a rotating polishing pad. A carrier applies pressure against the backside of the substrate. During the polishing process, the pad and table are rotated while a downward force is maintained against the substrate back. An abrasive and chemically reactive solution, commonly referred to as a "slurry" is applied to the pad during polishing. The slurry initiates the polishing process by chemically reacting with the film being polished. The polishing process is facilitated by the rotational movement of the pad relative to the substrate as slurry is provided to the wafer/pad interface. Polishing is continued in this manner until the desired film on the insulator is removed. The slurry composition is an important factor in the CMP step. Depending on the choice of the oxidizing agent, the abrasive, and other useful additives, the polishing slurry can be tailored to provide effective polishing to metal layers at desired polishing rates while minimizing surface imperfections, defects and corrosion and erosion. Furthermore, the polishing slurry may be used to provide controlled polishing selectivities to other thin-film materials used in current integrated circuit technology such as titanium, titanium nitride and the like.

Typically CMP polishing slurries contain an abrasive material, such as silica or alumina, suspended in an oxidizing, aqueous medium. For example, U.S. patent No. 5,244,523 to Yu et al. reports a slurry containing alumina, hydrogen peroxide, and either potassium or ammonium hydroxide that is useful to remove tungsten at predictable rates with little removal of the underlying insulating layer. U.S. Patent 5,209,816 to Yu et al. discloses a slurry comprising perchloric acid, hydrogen peroxide and a solid abrasive material in an aqueous medium that is useful for polishing aluminum. U.S. Patent 5,340,370 to Cadieu and Feller discloses a tungsten polishing slurry comprising approximately 0.1M potassium ferricyanide, approximately 5 weight percent silica and potassium acetate. Acetic acid is added to buffer the pH at approximately 3.5.

U.S. Patent No. 4,789,648 to Beyer et al. discloses a slurry formulation using alumina abrasives in conjunction with sulfuric, nitric, and acetic acids and deionized water. U.S. Patent Nos. 5,391,258 and 5,476,606 disclose slurries

for polishing a composite of metal and silica which includes an aqueous medium, abrasive particles and an anion which controls the rate of silica removal. Other polishing slurries for use in CMP applications are described in U.S. Patent No. 5,527,423 to Neville et al., U.S. Patent No. 5,354,490 to Yu et al., U.S. Patent No. 5,340,370 to Cadien et al., U.S. Patent No. 5,209,816 to Yu et al., U.S. Patent No. 5,157,876 to Medellin, U.S. Patent No. 5,137,544 to Medellin, and U.S. Patent No. 4,956,313 to Coto et al.

There are various mechanisms disclosed in the prior art by which metal surfaces can be polished with slurries. The metal surface may be polished using a slurry in which a surface film is not formed in which case the process proceeds by mechanical removal of metal particles and their dissolution in the slurry. In such a mechanism, the chemical dissolution rate should be slow in order to avoid wet etching. A more preferred mechanism is, however, one where a thin abrasable layer is continuously formed by reaction between the metal surface and one or more components in the slurry such as a complexing agent and/or a film forming layer. The thin abrasable layer is then removed in a controlled manner by mechanical action. Once the mechanical polishing process has stopped a thin passive film remains on the surface and controls the wet etching process. Controlling the chemical mechanical polishing process is much easier when a CMP slurry polishes using this mechanism.

Efforts to develop copper CMP slurries are disclosed in the literature. The RPI effort (J. M. Steigerwald et al, *Electrochemical Potential Measurements during the Chemical-Mechanical Polishing of Copper Thin Films*, Mat. Res. Soc. Symp. 337, 193 (1994)) is focused on the use of ammonium compounds (ammonium nitrate, chloride, hydroxide), nitric acid, and alumina abrasive. Copper dissolution of 2 nm/min (as measured electrochemically) is assumed to proceed from a film-free surface. Polishing rates, however, are reported to be in excess of 400 nm/min. The discrepancy is explained by importance given to the mechanical action, forming Cu debris, which is then dissolved by solution. Selectivity factors are not given.

Q. Luo et al, *Chemical-Mechanical Polishing of Copper in Acidic Media*, Proceedings - First International Chemical-Mechanical Polish (CMP) for VLSI/VLSI Multilevel Interconnection Conference (CMP-MIC), Santa Barbara, Feb. 22-23, (1996) discloses using a CMP slurry including a very aggressive etchant, Fe-nitrate, pH 1-2, in combination with an inhibitor benzotriazole, a slurry stabilizing surfactant (poly-ethyleneglycol) and alumina. The chemical reaction is apparently controlled by a formation of a corrosion inhibiting film, namely Cu-BTA, with surfactant undermining its protectiveness. Selectivity to oxide is given as 15:1 to 45:1.

CMP electrochemical work at Sematech laboratories is disclosed in R. Carpio et al, *Initial Study On Copper CMP Slurry Chemistries*, Thin Solid Films, 262 (1995). The reference explores the use of electrochemistry in the fundamental characterization of plausible slurries. In addition to several others, potassium permanganate is used as a slurry oxidizer.

H. Hirabayashi et al, *Chemical Mechanical Polishing of Copper Using A Slurry Composed of Glycine and Hydrogen Peroxide*, Proceedings - First International Chemical-Mechanical Polish (CMP) for VLSI/VLSI Multilevel Interconnection Conference (CMP-MIC), Santa Barbara, Feb. 22-23, (1996), and Japanese Kokai Patent Application No. 8 (1996) 83780 disclose a mixture of glycine, hydrogen peroxide and silica, with or without benzotriazole, for the CMP process of Cu with a low corrosion rate and defect level. The references disclose that CMP slurries incorporating a chemical agent, such as benzotriazole and n-benzoyl-n-phenylhydroxylamine form a protective film on copper. The removal rate varies, depending on the concentration of slurry components. An optimized rate of 120 nm/min was reported, with TIN rate of 30 nm/min and dishing of 200 nm across the 15 μ m wide structures.

Several relevant Cu chemistries have been discussed in the open literature, each failing to deliver a process which successfully addresses all of the key requirements of a chemical-mechanical polishing slurry; namely metal removal rate of more than 200 nm/min, rate selectivity to metal lines of <5, selectivity to dielectric oxide layer of >50 and overall defect depth of <10%.

Despite the desirability of using a film forming mechanism in a CMP process there remains problems with formulating CMP slurries that can control the thickness of the layer of film formed as well as problems ensuring that the film formed is abrasable. These problems can result in a CMP slurry that exhibits unacceptably low polishing rates or poor polishing results. Thus, a need remains for a CMP slurry that is capable of forming a removable thin abrasable layer on a substrate surface and more particularly on the surface of a copper alloy containing substrate. A desirable CMP slurry will exhibit good thin film polishing selectivities and simultaneously give polished substrates with minimal dishing and low defectivity.

SUMMARY OF THE INVENTION

The present invention is directed to a chemical mechanical polishing slurry that is able to polish metal containing substrates at acceptable rates.

In addition, the chemical mechanical polishing slurry has a low insulator polishing selectivity while exhibiting high polishing selectivities towards copper and copper alloy containing metal layers.

Furthermore, this invention is directed to methods for using a single chemical mechanical polishing slurry to polish metal layers and particularly copper or copper alloy containing layers in an integrated circuit.

This invention is also a chemical mechanical polishing slurry that includes urea hydrogen peroxide as a stable effective oxidizing agent.

In one embodiment, this invention is a chemical mechanical polishing slurry. The CMP slurry includes an abrasive, an oxidizer, a complexing agent, and a film forming agent. In one preferred embodiment, the CMP slurry includes from about 1.0 to about 15.0 weight percent of an alumina abrasive, from about 0.3 to about 12.0 weight percent hydrogen peroxide, from about 0.5 to about 3.0 weight percent ammonium oxalate or tartaric acid, and from about 0.01 to about 0.2 weight percent benzotriazole. In another preferred embodiment the CMP slurry includes 1.0 to about 15.0 wt % alumina abrasive, from about 1.0 to 12.0 wt % urea hydrogen peroxide, from 0.5 to 3.0 wt % ammonium oxalate or tartaric acid and from about 0.01 to about 0.2 wt % benzotriazole. All CMP slurry embodiments include deionized water as well.

In another embodiment, this invention is a method for polishing a substrate including at least one metal layer. The method begins by mixing from about 1.0 to about 15.0 weight percent of an abrasive, from about 0.3 to about 12.0 weight percent of an oxidizer and preferably with urea hydrogen peroxide, from about 0.5 to about 3.0 weight percent of at least one complexing agent, from about 0.01 to about 0.2 weight percent of at least one film forming agent, and deionized water to give a chemical mechanical polishing slurry. Next, the chemical mechanical polishing slurry is applied to the substrate. Finally, at least a portion of the metal layer is removed from the substrate by bringing a pad into contact with the substrate and moving the pad in relation to the substrate.

In still another embodiment, this invention is a CMP slurry precursor comprising urea and at least one metal oxide abrasive. The CMP slurry precursor is combined with hydrogen peroxide in a mole ratio of urea to hydrogen peroxide of from about 0.75:1 to about 2:1 prior to use to give a CMP slurry.

DESCRIPTION OF THE FIGURES

Figure 1 shows potentiodynamic polarization curves measured on copper in solutions containing 4 weight % of ammonium persulfate and either 1 weight % glycine (curves 1 & 2) or 1 weight % ammonium oxalate (curves 3 & 4). Each set of curves was measured during copper surface abrasion (curves 1 & 3), and again after abrasion had ceased (curves 2 & 4).

Figure 2 shows potentiodynamic polarization curves on copper in: 11% H_2O_2 wt % oxidizer solution (curve 1 after abrasion); in electrolyte with the same oxidizer and 1 wt % glycine (curve 2, after abrasion); and in the same oxidizer with 1 wt % ammonium oxalate (curve 3, also after abrasion).

Figure 3 shows the reproducibility of the polishing performance, i.e., polishing rate and within wafer non-uniformity, for copper using a CMP slurry of this invention comprising 5.0 wt % alumina abrasive, 11.0 wt % H_2O_2 , 1.5 wt % ammonium oxalate, 0.04 wt % benzotriazole, and 50ppm wetting agent (TRITON® DF-16).

DESCRIPTION OF THE CURRENT EMBODIMENT

The present invention relates to a chemical mechanical polishing slurry that comprises an abrasive, at least one oxidizer, a complexing agent and a film forming agent. The chemical mechanical polishing slurry is useful for polishing metals, especially copper and copper alloy containing metal layers associated with a substrate selected from the group including integrated circuits, thin films, multiple level semiconductors, and wafers.

Before describing the details of the various preferred embodiments of this invention, some of the terms that are used herein will be defined. The chemical mechanical polishing slurry, ("CMP slurry"), is a useful product of this invention that comprises an oxidizer, an abrasive, a complexing agent, a film forming agent, and other optional ingredients. The CMP slurry is useful for polishing a multiple level metallization which may include but are not limited to semiconductor thin-films, integrated circuit thin-films, and for any other films and surfaces where CMP processes are useful. The terms "copper" and "copper containing alloys" are used interchangeably herein as it is within the understanding of one of skill in the art that the terms include but are not limited to substrates comprising layers of pure copper, copper aluminum alloys, and Ti/TiN/Cu, and Ta/TaN/Cu multi-layer substrates.

The CMP slurry of this invention includes at least one oxidizer. The oxidizer aids in oxidizing the substrate metal layer or layers to their corresponding oxide, hydroxide, or ions. For example, in the present invention, the oxidizer may be used to oxidize a metal layer to its corresponding oxide or hydroxide, e.g., titanium to titanium oxide, tungsten to tungsten oxide, copper to copper oxide, and aluminum to aluminum oxide. The oxidizing agent is useful when incorporated into a CMP slurry to polish metals and metal based components including titanium, titanium nitride, tantalum, copper, tungsten, aluminum, and aluminum alloys such as aluminum/copper alloys, and various mixtures and combinations thereof by mechanically polishing the metals to remove the respective oxide layer.

The oxidizer used in the CMP slurry of this invention may be selected from compounds which, upon reduction, form hydroxyl radicals. Such oxidizers exhibit good polishing selectivity towards metal and metal containing substrate layers and particularly towards copper alloy layers. Non-exclusive examples of metal oxidizing compounds that, upon

reduction, form hydroxyl radicals include peracetic acid, urea-hydrogen peroxide, urea peroxide, and hydrogen peroxide and mixtures thereof, with hydrogen peroxide and urea hydrogen peroxide being preferred oxidizers. The oxidizer may be present in the chemical mechanical polishing slurry in an amount ranging from about 0.3 to about 30.0 weight percent. It is preferred that the oxidizer is present in the CMP slurry of this invention in an amount ranging from about 0.3 to about 17.0 weight percent and most preferably from about 1.0 to about 12.0 weight percent.

A preferred oxidizer is urea hydrogen peroxide. Because urea hydrogen peroxide is 34.5 wt % hydrogen peroxide and 65.5 wt % urea, a greater amount by weight of urea hydrogen peroxide must be included in CMP slurries of this invention to achieve the desired oxidizer loading set forth above. For example, a range of 1.0 to 12.0 weight percent oxidizer corresponds to a urea hydrogen peroxide weight three times as great or from 3.0 to 36.0 weight percent.

A CMP slurry comprising urea hydrogen peroxide can be formulated by a number of methods including combining urea peroxide with water, and by combining urea and hydrogen peroxide in an aqueous solution in a mole ratio range of from about 0.75:1 to about 2:1 to give a urea hydrogen peroxide oxidizer.

The CMP slurry of this invention also includes a film forming agent. The film forming agent may be any compound or mixtures of compounds that are capable of facilitating the formation of a passivation layer of metal oxides and dissolution inhibiting layers on the surface of the metal layer. Passivation of the substrate surface layer is important to prevent wet etching of the substrate surface. Useful film forming agents are cyclic compounds such as imidazole, benzotriazole, benzimidazole and benzothiazole and their derivatives with hydroxy, amino, imino, carboxy, mercapto, nitro and alkyl substituted groups, as well as urea, thiourea and others. A preferred film forming agent is benzotriazole ("BTA"). The film forming agent should be present in the chemical mechanical polishing slurry of this invention in an amount that is capable of promoting quick, and preferably almost instantaneous formation of passivating layers and dissolution inhibiting layers on the substrate surface. The film forming agent should be present in the CMP slurry of this invention in an amount ranging from about 0.01 weight percent to about 1.0 weight percent. It is preferred that film forming agent is present in the CMP slurry in an amount ranging from about 0.01 to about 0.2 weight percent.

Once a passivation layer has formed on the substrate surface it becomes important to be able to disturb the passivation layer in order to abrade metal oxides from the substrate surface with the abrasive component of the CMP slurry of this invention. One class of compounds that is useful in disturbing the passivation layer are complexing agents. Useful complexing agents include but are not limited to acids such as citric, lactic, tartaric, succinic, acetic, oxalic and other acids, as well as amino acid and amino sulfuric acids and their salts. A preferred complexing agent is ammonium oxalate. Another preferred complexing agent is tartaric acid.

The complexing agents serve at least two useful functions in the CMP slurry of this invention. The complexing agent disturbs the passivation layer during the mechanical abrasion step without destroying the layer or inhibiting its formation during the abrasion step and especially after the abrasion step is complete. Secondly, the complexing agent is believed to form a complex with the oxidized metal and not the underlying unoxidized metal thereby limiting the depth of the oxidized layer. The complexing agent will be present in the CMP slurry of this invention in an amount ranging from about 0.5 to about 5.0 weight percent and preferably in an amount ranging from about 0.5 to about 3.0 weight percent.

Other well known polishing slurry additives may be incorporated into the chemical mechanical polishing slurry of this invention. One class of optional additives are inorganic acids and/or salts thereof which may be added to the polishing slurry to further improve or enhance the polishing rate of the barrier layers in the wafer, such as titanium and tantalum. Useful inorganic additives include sulfuric acid, phosphoric acid, nitric acid, HF acid, ammonium fluoride, ammonium salts, potassium salts, sodium salts or other cationic salts of sulfates, phosphates and fluorides.

BTA, or other film forming agents in the CMP slurry of this invention may destabilize the uniform dispersion of abrasive in the slurry. In order to promote stabilization of a CMP slurry of this invention against settling, flocculation, and decomposition, a variety of optional CMP slurry additives, such as surfactants, stabilizers, or dispersing agents, can be used. If a surfactant is added to the CMP slurry, then it may be an anionic, cationic, nonionic, or amphoteric surfactant or a combination of two or more surfactants can be employed. Furthermore, it has been found that the addition of a surfactant may be useful to reduce the within-wafer-non-uniformity (WIWNU) of the wafers, thereby improving the surface of the wafer and reducing wafer defects.

In general, the amount of additive such as a surfactant that may be used in the present invention should be sufficient to achieve effective stabilization of the slurry and will typically vary depending on the particular surfactant selected and the nature of the surface of the metal oxide abrasive. For example, if not enough of a selected surfactant is used, it will have little or no effect on CMP slurry stabilization. On the other hand, too much surfactant in the CMP slurry may result in undesirable foaming and/or flocculation in the slurry. As a result, stabilizers such as surfactants should generally be present in the slurry of this invention in an amount ranging from about 0.001% to about 0.2% by weight, and preferably from about 0.001 to about 0.1 weight percent. Furthermore, the additive may be added directly to the slurry or treated onto the surface of the metal oxide abrasive utilizing known techniques. In either case, the amount of additive is adjusted to achieve the desired concentration in the polishing slurry. Preferred surfactants include dodecyl sulfate sodium salt, sodium lauryl sulfate, dodecyl sulfate ammonium salt, and mixtures thereof. Examples of useful surfactants

include TRITON® DF-16 manufactured by Union Carbide, and SURFYNOL® manufactured by Air Products and Chemicals.

The CMP slurry of this invention includes an abrasive. The abrasive is typically a metal oxide abrasive. The metal oxide abrasive may be selected from the group including alumina, titania, zirconia, germania, silica, ceria and mixtures thereof. The CMP slurry of this invention preferably includes from about 1.0 to about 15.0 weight percent or more of an abrasive. It is more preferred, however, that the CMP slurry of this invention includes from about 3.0 to about 6.0 weight percent abrasive.

The metal oxide abrasive may be produced by any techniques known to those skilled in the art. Metal oxide abrasives can be produced using any high temperature process such as sol-gel, hydrothermal or, plasma process, or by processes for manufacturing fumed or precipitated metal oxides. Preferably, the metal oxide is a fumed or precipitated abrasive and, more preferably it is a fumed abrasive such as fumed silica or fumed alumina. For example, the production of fumed metal oxides is a well-known process which involves the hydrolysis of suitable feedstock vapor (such as aluminum chloride for an alumina abrasive) in a flame of hydrogen and oxygen. Molten particles of roughly spherical shapes are formed in the combustion process, the diameters of which are varied through process parameters. These molten spheres of alumina or similar oxide, typically referred to as primary particles, fuse with one another by undergoing collisions at their contact points to form branched, three dimensional chain-like aggregates. The force necessary to break aggregates is considerable. During cooling and collecting, the aggregates undergo further collision that may result in some mechanical entanglement to form agglomerates. Agglomerates are thought to be loosely held together by van der Waals forces and can be reversed, i.e., de-agglomerated by proper dispersion in a suitable media.

Precipitated abrasives may be manufactured by conventional techniques such as by coagulation of the desired particles from an aqueous medium under the influence of high salt concentrations, acids or other coagulants. The particles are filtered, washed, dried and separated from residues of other reaction products by conventional techniques known to those skilled in the art.

A preferred metal oxide will have a surface area, as calculated from the method of S. Brunauer, P.H. Emmet, and I. Teller, J. Am. Chemical Society, Volume 60, Page 309 (1938) and commonly referred to as BET, ranging from about 5 m²/g to about 430 m²/g and preferably from about 30 m²/g to about 170 m²/g. Due to stringent purity requirements in the IC industry the preferred metal oxide should be of a high purity. High purity means that the total impurity content, from sources such as raw material impurities and trace processing contaminants, is typically less than 1% and preferably less than 0.01% (i.e., 100 ppm).

The metal oxide abrasive useful in the dispersion of this invention may consist of metal oxide aggregates or individual single sphere particles. The term "particle" as it is used herein refers to both aggregates of more than one primary particle and to single particles.

It is preferred that the metal oxide abrasive consists of metal oxide particles having a size distribution less than about 1.0 micron, a mean particle diameter less than about 0.4 micron and a force sufficient to repel and overcome the van der Waals forces between abrasive aggregates themselves. Such metal oxide abrasive has been found to be effective in minimizing or avoiding scratching, pit marks, divots and other surface imperfections during polishing. The particle size distribution in the present invention may be determined utilizing known techniques such as transmission electron microscopy (TEM). The mean particle diameter refers to the average equivalent spherical diameter when using TEM image analysis, i.e., based on the cross-sectional area of the particle. By force is meant that either the surface potential or the hydration force of the metal oxide particles must be sufficient to repel and overcome the van der Waals attractive forces between the particles.

In another preferred embodiment, the metal oxide abrasive may consist of discrete, individual metal oxide particles having a primary particle diameter less than 0.4 micron (400nm) and a surface area ranging from about 10 m²/g to about 250 m²/g.

Preferably, the metal oxide abrasive is incorporated into the aqueous medium of the polishing slurry as a concentrated aqueous dispersion of metal oxides, comprising from about 3% to about 45% solids, and preferably between 10% and 20% solids. The aqueous dispersion of metal oxides may be produced utilizing conventional techniques, such as slowly adding the metal oxide abrasive to an appropriate media, for example, deionized water, to form a colloidal dispersion. The dispersion is typically completed by subjecting it to high shear mixing conditions known to those skilled in the art. The pH of the slurry may be adjusted away from the isoelectric point to maximize colloidal stability.

It is desirable to maintain the pH of the CMP slurry of this invention within a range of from about 2.0 to about 12.0, preferably between from about 4.0 to about 9.0 and most preferably from about 5.0 to about 8.0 in order to facilitate control of the CMP process. Slurry handling problems and substrate polishing quality problems are encountered when the pH of the CMP slurry of this invention is too low, e.g., less than 2. When ammonium oxalate is the complexing agent the CMP precursor or slurry will typically have a pH of about 7.5 so no pH adjustment may be necessary. However, when tartaric acid is selected as the complexing agent, the CMP precursor or slurry will typically have a pH of about 2.0 and pH adjustment is preferred.

The pH of the CMP precursors and slurries of this invention may be adjusted using any known acid, base, or amine.

However, the use of an acid or base that contains no metal ions, such as ammonium hydroxide and amines, or nitric, phosphoric, sulfuric, or organic acids are preferred to avoid introducing undesirable metal components into the CMP slurry of this invention.

Although the CMP slurry of this invention may be used to polish any type of metal layer, the chemical mechanical polishing slurry of this invention has been found to have a high copper, titanium, titanium nitride, and tantalum nitrate and acceptable tantalum polishing rates. In addition, the chemical mechanical polishing slurry exhibits desirable low polishing rates towards the dielectric insulating layer.

The CMP slurry of this invention may be produced using conventional techniques known to those skilled in the art. Typically, the oxidizing agent and other non-abrasive components, are mixed into an aqueous medium, such as deionized or distilled water, at predetermined concentrations under low shear conditions until such components are completely dissolved in the medium. A concentrated dispersion of the metal oxide abrasive, such as fumed alumina, is added to the medium and diluted to the desired loading level of abrasive in the final CMP slurry.

The CMP slurries of the present invention may be supplied as one package system (oxidizing agent, abrasive, film forming agent and passivating agent in a stable aqueous medium). To avoid possible CMP slurry degradation, however, it is preferred that at least a two package system is used where the first package comprises the film forming agent and any optional additives, and the second package comprises the aqueous abrasive dispersion and an oxidizer. The remaining component, the complexing agent, may be placed in either the first container, the second container or in a third container. Other two-container combinations of the ingredients of the CMP slurry of this invention are within the knowledge of one having ordinary skill in the art.

It has been discovered that CMP slurries of this invention including urea hydrogen peroxide can be formulated by adding hydrogen peroxide to a slurry precursor comprising urea and any other useful slurry components to give a urea hydrogen peroxide containing CMP slurry. Formulating CMP slurries of this invention from a urea containing slurry precursor eliminates stability, shipping and safety concerns associated with hydrogen peroxide containing slurries. This is because the urea containing CMP slurry precursor can be prepared and shipped to the location where it will be used and then mixed with hydrogen peroxide available on site to give a CMP slurry including urea hydrogen peroxide.

A preferred slurry precursor of this invention will comprise a dry or aqueous mixture of urea and at least one metal oxide abrasive. Additional ingredients may be incorporated into the urea containing slurry precursor including at least one complexing agent, at least one film forming agent, and any other additives such as surfactants that are useful in CMP slurries.

A most preferred slurry precursor of this invention includes an aqueous dispersion of from about 2.0 to about 24.0 wt % urea, fumed alumina, a complexing agent selected from ammonium oxalate, tartaric acid or mixtures thereof, and preferably tartaric acid, benzotriazole, and a surfactant in quantities disclosed above. The slurry precursor or mixtures thereof, will preferably have a pH of from about 4.0 to about 9.0.

A multi-package CMP slurry system may be used with any standard polishing equipment appropriate for use on the desired metal layer of the wafer. The multi-package system includes one or more CMP slurry components in, where appropriate, aqueous or dry form in two or more containers. The multi-package system is used by combining the components from the various containers in the desired amounts to give a CMP slurry comprising at least one oxidizing agent, a film forming agent, a complexing agent and at least one abrasive in the amounts described above prior to or at the time of the slurry to a substrate. The preferred package system comprises a first container including a CMP slurry precursor comprising alumina, urea, a complexing agent selected from ammonium oxalate, tartaric acid, and mixtures thereof, benzotriazole, and a surfactant at a pH from about 4.0 to about 9.0 and a second container including hydrogen peroxide. At the polishing location, a preselected amount of the CMP precursor and a selected amount of hydrogen peroxide are combined at the time of polishing to give a CMP slurry of this invention.

The CMP slurry of the present invention does not significantly increase the silicon dioxide polishing rate. However, the CMP slurry of this invention polishes copper, titanium, titanium nitride, tantalum, and tantalum nitride layers at good rates under controllable conditions. Thus, the CMP slurry of this invention is effective in controlling polishing selectivities of titanium, copper, and titanium nitride. The polishing slurry of the present invention may be used during the various stages of semiconductor integrated circuit manufacture to provide effective polishing at desired polishing rates while minimizing surface imperfection and defects.

EXAMPLES

We have discovered that a CMP slurry including at least one oxidizer, a complexing agent and a film forming agent is capable of polishing multiple metal layers comprising copper alloys, titanium, and titanium nitride at high rates while exhibiting an acceptable low polishing rate towards dielectric layers.

The following examples illustrate preferred embodiments of this invention as well as preferred methods for using compositions of this invention.

EXAMPLE 1

Electrochemical tests are used to evaluate CMP mechanisms and to provide guidance in selection of slurry components. The electrochemical cell used was developed in the IBM Laboratories and is described in V. Brusic et al., *Corrosion And Inhibition Of Thin Line Conductors In VLSI Structures*, IBM J R&D, 37, 173 (1993), incorporated herein by reference. The electrochemical cell allows for the evaluation of the electrode potential and the metal dissolution rate of a substrate, with and without surface abrasion. The method uses a PAR model 273 potentiostat with PAR corrosion software.

Figure 1 illustrates processes governing the dissolution of copper in the presence of an oxidizer, i.e., 4 weight % ammonium persulfate with either glycine (1 & 2) or ammonium oxalate (3 & 4), both present in the amount of one percent by weight at pH 5.1. In both cases the anodic Tafel slope is quite low, and the difference between the potentiodynamic polarization curves with and without abrasion is very small. The electrochemical cell results indicate that the Cu surface is not covered by an oxide film. However, copper dissolution is at least twenty times slower in the presence of ammonium oxalate. In a comparison to glycine, oxalate is a more effective inhibitor to copper dissolution. There is also a persistent difference in the corrosion potentials, with the potential in the ammonium oxalate being consistently lower than the potential measured in glycine, indicating that a preferential dissolution process would lead to Cu^+ ions while with glycine, Cu^{++} ion formation is possible.

At higher pHs, oxalate still adsorbs at the copper surface, but ammonium oxalate also acts to increase the copper dissolution rate through formation of $\text{Cu}(\text{NH}_3)_4^{++}$ and $\text{Cu}(\text{NH}_3)_5^{+}$ complexes. Figure 2 shows the dissolution and passivation of copper in 11% H_2O_2 (curve 1, after surface abrasion), the same oxidizer with 1 weight % glycine (curve 2) and the same oxidizer with 1% ammonium oxalate (curve 3). Corrosion potential is lowest in the presence of ammonium oxalate. Although in ammonium oxalate the copper dissolution is enhanced in a comparison to peroxide alone, reaching about 200 nm/min during abrasion, the surface repassivation does occur, leading to the dissolution rate after abrasion of only 5.5 nm/min. The additions of small amounts of BTA assure that the passivation should occur promptly, with Cu-BTA providing an additional factor in copper dissolution control. In contrast, copper dissolution in glycine is the same with and without abrasion, reaching the uncontrollable values of over 300 nm/min without surface repassivation.

EXAMPLE 2

Cu and Ti wafers were polished with a CMP slurry using a Strasbaugh polisher at 3 psi down force, 45 rpm table speed, and 50 rpm spindle speed. The CMP slurry was formulated to vary the concentration of hydrogen peroxide, ammonium oxalate, benzotriazole and wetting agent as shown in Table 1. Cu and Ti removal rates were measured. Several rates for prior art slurries including hydrogen peroxide and glycine were also determined and reported in Table 2. All slurries contained alumina abrasive, with 5% solids. The wetting agent used was TRITON® DF16 manufactured by Union Carbide Chemicals & Plastics Co., Danbury, Connecticut.

Table 1

Polishing rates of Cu and Ti in peroxide slurries					
Peroxide (H_2O_2)	Ammonium Oxalate	BTA	Wetting Agent	Cu Rate nm/min	Ti Rate nm/min
7%	0%	0	50 ppm	21.7	80.6
7%	0.5%	0	50 ppm	278	30.7
11%	0.5%	0	10 ppm	251.7	25.4
11%	1.0%	0	10 ppm	402.9	80.4
9%	1.0%	0.04%	30 ppm	170.7	94.1
7%	1.5%	0.08%	10 ppm	304.7	108.6

Table 2

Peroxide (H_2O_2)	Glycine	BTA	Wetting Agent	Cu Rate nm/min	Ti Rate nm/min
11%	0.1%	0	0	52.8	101.4
8%	1.1%	0	25 ppm	493.7	75.6
11%	2.1%	0	0	778.3	53.4

EP 0 846 742 A2

The results of the CMP polishing tests set forth in Tables 1 and 2 show that CMP slurries of this invention are capable of achieving preferred copper and titanium polishing rates and selectivities including a Cu rate of at least 100 nm/min and a [Cu:Ti] selectivity of at most 4:1.

EXAMPLE 3

The reproducibility of Cu removal rates and within wafer non-uniformity (WIWNU) results for a CMP slurry containing 11.0 weight hydrogen peroxide, 1.5% ammonium oxalate, 0.04% BTA, 50 ppm of TRITON® DF-16 surfactant manufactured by Union Carbide, and 5% alumina abrasive was evaluated in this Example. The CMP slurry was applied successively to copper wafers on a Strasbaugh polisher, using Rodel 1158 buffed pad with DF 200 insert, a 4 psi down force, a table speed of 50 rpm, and a spindle speed of 50 rpm.

The experimental results, plotted in Figure 3 demonstrate that the slurry polishing performance is quite consistent and shows good wafer to wafer uniformity.

EXAMPLE 4

This Example compares the effectiveness of urea hydrogen peroxide and hydrogen peroxide as useful CMP oxidizers. Specifically, this Example compares the stability over time of the two oxidizers.

Two slurries having the following compositions were prepared in an aqueous medium (deionized water). Each slurry was prepared using SEMI-SPERSE® W-A335 alumina dispersion and diluted to 5 weight percent alumina with deionized water.

Slurry A: 5% alumina, 3% hydrogen peroxide [HP], 3% succinic acid, original pH = 3.50.

Slurry B: 5% alumina, 8.5% urea-hydrogen peroxide [UHP] (which corresponds to about 3.0 weight percent aqueous H₂O₂), 3% succinic acid, original pH = 3.55.

Slurries A and B were allowed to stand at room temperature, over a seven week period. Samples of slurry A and B were periodically analyzed for pH and titrated with potassium permanganate in acid solution to determine percentage of active peroxide. The results are reported in Table 3 below.

Table 3 -

Stability Study of Slurries A and B				
	Active H ₂ O ₂ %		pH	
	Slurry A	Slurry B	Slurry A	Slurry B
	HP	UHP	HP	UHP
0 weeks	3.35	3.28	3.50	3.55
1 weeks		3.30		3.48
2 weeks	2.85		3.62	
3 weeks		3.24		3.52
5 weeks		3.12		3.49
7 weeks	1.82		3.54	
Average change/week	-0.22	-0.03	0.006	-0.012

The test results indicate that the active peroxide in the slurry including hydrogen peroxide degrades at a much faster rate than the slurry including urea hydrogen peroxide. The pH stability of both slurries are similar.

A most preferred oxidizer is urea hydrogen peroxide. The urea hydrogen peroxide may be present in the overall chemical mechanical polishing slurry in an amount ranging from about 1.5 to about 30.0 weight percent. It is preferred that urea hydrogen peroxide is present in the slurry in an amount ranging from about 3.0 to about 17.0 weight percent and most preferably from about 5.0 to about 12.0 weight percent.

EXAMPLE 5

The effectiveness of polishing slurries comprising urea hydrogen peroxide to polish Cu and Ti wafers was evaluated in this Example. The slurries described in Table 4 below were applied to Cu and Ti wafers on a IPEC 472 polisher using a IC1000/SUBA IV pad stack manufactured by Rodel, Inc. at 5 psi down force, a table speed of 50 rpm, and a spindle

EP 0 846 742 A2

speed of 60 rpm. Each slurry included 5.0 wt % alumina abrasive. The wetting agent used was TRITON® DF-16. The Cu and Ti polishing rates are reported in Table 4.

Table 4

Polishing rates of Cu and Ti in urea-peroxide slurries					
Urea-peroxide	Ammonium oxalate	BTA	Wetting agent	Cu rate nm/min	Ti rate nm/min
5.65%	0%	0%	0%	9.5	260.7
5.65%	1%	0%	50 ppm	796.3	249.4
5.65%	1%	0.08%	50 ppm	673.5	271.0

The polishing results indicate that CMP slurries of this invention are capable of achieving preferred Cu and Ti polishing rates of over 500nm/min and a Cu:Ti selectivity ratio of less than 2.5.

EXAMPLE 6

This example evaluated the effectiveness of CMP polishing slurries comprising a polishing slurry precursor and hydrogen peroxide copper polishing. The slurries described in Table 5, below, were prepared by combining an aqueous dispersion of a CMP slurry precursor consisting of 5 wt % alumina, urea, ammonium oxalate, BTA, and wetting agent (TRITON®DF-16) with a 30 wt % solution of H₂O₂. The resulting CMP slurries were applied to Cu wafers on a IPEC 472 polisher using a IC1000/SUBA 500 pad stack manufactured by Rodel, Inc., at 5 psi down force, a table speed of 40 rpm and a spindle speed of 60 rpm. The Cu polishing rates are reported in Table 5, below.

Table 5

Urea	Ammonium Oxalate	BTA	Wetting agent	Hydrogen peroxide (30%)	Cu rate nm/min
0%	1%	0.08%	50 ppm	8.87%	572.7
3.65%	1%	0.08%	50 ppm	6.67%	508.3
7.3%	1%	0.08%	50 ppm	6.67%	506.5

Results indicate that CMP slurries of this invention are capable of achieving preferred Cu (and Ti) polishing rates of over 500 nm/min and Cu:Ti selectivity ratio of less than 2.5. Useful slurries may be prepared either with a solid urea-peroxide as an oxidizer as shown in Table 4 of Example 5 or by combining liquid hydrogen peroxide (packaged as 30% solution), with a CMP slurry precursor comprising solid urea as shown in Table 5.

EXAMPLE 7

This example evaluated the effectiveness of CMP polishing slurries comprising tartaric acid and ammonium oxalate on copper, tantalum, and PTEOS polishing. Two slurries were prepared by combining ingredients to give a CMP slurry including 3 wt% alumina 3.85 wt% urea, 2.0 wt% hydrogen peroxide, 50 ppm Triton DF-16 surfactant, and 0.04 wt% benzotriazole. Slurry 1 included 1.0 wt% ammonium oxalate while slurry 2 included 3.0 wt% tartaric acid. The pH of the ammonium oxalate containing slurry was naturally 7.5, the pH of the tartaric acid containing slurry was adjusted to 7.5 by the addition of ammonium hydroxide to the slurry.

Both CMP slurries were applied to Cu, Ta, and PTEOS wafers on a IPEC 472 polisher using a IC1000/SUBA 500 pad stack manufactured by Rodel, Inc., at 3 psi down force, a 2 psi back pressure, a table speed of 55 rpm and a spindle speed of 30 rpm. The polishing rates are reported in Table 6, below.

TABLE 6

Slurry	Cu Removal Rate (Å/min)	Ta Removal Rate (Å/min)	PTEOS Removal Rate (Å/min)
1	4485	345	80
2	3746	208	67

The polishing rates of the tartaric acid containing polishing slurry are slightly lower than the rates of the ammonium oxalate containing slurry. However, the tartaric acid containing CMP slurry is more passivating than the ammonium

oxalate containing slurry thereby providing stronger corrosion control.

EXAMPLE 8

Corrosion rate of Cu using five different slurries was evaluated electrochemically according to the method of Example 1. All five slurries contained 5% dispersed alumina, 2% H_2O_2 , 3.85% urea and 50 ppm Trilon DF-16 surfactant. The complexing agent, the presence or absence of an inhibitor and the slurry pH are listed in the Table 7, along with the measured corrosion rates.

TABLE 7

Slurry	complexing agent	BTA	pH	Corrosion rate A/min
1	1% amm.oxalate	0%	7.5 (natural)	48
2	1% amm.oxalate	0.4%	7.5 (natural)	1.54
3	1% tartaric acid	0%	7.5 (w/ NH_4OH)	20
4	1% tartaric acid	0.4%	7.5 (w/ NH_4OH)	0.6
5	1% tartaric acid	0.4%	7.5 (w/TMAH)	0.5

The corrosion data indicates that tartaric acid lowers the Cu corrosion rate in a comparison with ammonium oxalate. The corrosion rate is further reduced in the presence of tetra alkyl ammonium hydroxide.

Claims

1. A CMP slurry precursor composition comprising urea and at least one metal oxide abrasive.
2. The composition of claim 1, further including a film forming agent.
3. The composition of claims 1 or 2, further including a complexing agent.
4. The composition of claims 1-3, wherein said composition comprises from about 2.0 weight percent to about 24.0 weight percent urea and from about 1.0 to about 15.0 weight percent of an alumina abrasive.
5. The composition of claim 4, wherein said composition further comprises from about 0.5 to about 3.0 weight percent ammonium oxalate; and from about 0.01 to about 0.2 weight percent benzotriazole; and from about 0.001 to about 0.1 weight percent of a surfactant.
6. The composition of claim 4, wherein said composition further comprises from about 0.5 to about 5.0 weight percent tartaric acid and from about 0.1 to about 0.2 weight percent benzotriazole.
7. The composition of claims 1-4 or 6, wherein the composition has a pH from about 5 to about 8.
8. The composition of claims 1-4 or 6-7, wherein the composition further includes from about 0.001 to about 0.1 weight percent of a surfactant.
9. A chemical mechanical polishing slurry composition comprising:
 - an abrasive;
 - at least one oxidizer;
 - a complexing agent; and
 - a film forming agent.
10. A chemical mechanical polishing slurry composition comprising:
 - an abrasive;
 - urea hydrogen peroxide;

a complexing agent; and
a film forming agent.

- 5 11. The composition of claims 3-10, wherein the complexing agent is selected from the group of compounds including citric acid, lactic acid, tartaric acid, succinic acid, oxalic acids, amino acids and salts thereof.
12. The composition of claims 3-11, wherein the complexing agent is ammonium oxalate.
- 10 13. The composition of claims 12, wherein the ammonium oxalate is present in an amount ranging from 0.5 to about 9.0 weight percent.
14. A chemical mechanical polishing slurry composition comprising:
15 an abrasive;
urea hydrogen peroxide;
tartaric acid; and
a film forming agent.
- 20 15. The composition of claim 14, wherein the tartaric acid is present in an amount ranging from about 0.5 to about 5.0 weight percent.
16. The composition of claims 2-15, wherein the film forming agent is benzotriazole.
- 25 17. The composition of claim 16, wherein the benzotriazole is present in an amount between 0.01 to about 0.1 weight percent.
18. The composition of claim 9, wherein the oxidizer is a compound that forms hydroxyl radicals upon reduction.
- 30 19. The composition of claims 9 or 18, wherein the oxidizer is selected from the group consisting of hydrogen peroxide, urea hydrogen peroxide and combinations thereof.
20. The composition of claims 9-19, wherein the oxidizer is hydrogen peroxide or urea hydrogen peroxide present in an amount ranging from about 0.3 to about 12 weight percent.
- 35 21. The composition of claims 9-20, wherein the abrasive is at least one metal oxide.
22. The composition of claims 1-21, wherein the metal oxide abrasive is selected from the group including alumina, ceria, germania, silica, titania, zirconia, and mixtures thereof.
- 40 23. The composition of claims 1-22, wherein the abrasive is an aqueous dispersion of a metal oxide.
24. The composition of claims 1-23, wherein the metal oxide abrasive consists of metal oxide aggregates having a size distribution less than about 1.0 micron and a mean aggregate diameter less than about 0.4 micron.
- 45 25. The composition of claims 1-23, wherein the metal oxide abrasive consists of discrete, individual metal oxide spheres having a primary particle diameter less than 0.400 micron and a surface area ranging from about 10 m²/g to about 250 m²/g.
- 50 26. The composition of claims 1-25, wherein the abrasive has a surface area ranging from about 5 m²/g to about 430 m²/g.
27. The composition of claim 26, wherein the abrasive has a surface area of from about 30 m²/g to about 170 m²/g.
- 55 28. The composition of claims 1-27, wherein the abrasive is selected from the group consisting of precipitated abrasives or turned abrasives.
29. The composition of claims 1-28, wherein the composition has a pH of from about 2 to about 12.

EP 0 846 742 A2

30. The composition of claim 29, wherein the composition has a pH of from about 4 to about 9.

31. The composition of claim 30, wherein the composition has a pH of from about 5 to about 8.

5 32. A chemical mechanical polishing slurry composition comprising:

an abrasive;
an oxidizing agent selected from the group consisting of hydrogen peroxide,
urea hydrogen peroxide, and mixtures thereof;
10 ammonium oxalate; and
benzotriazole.

33. The composition of claim 32, wherein the composition comprises:

15 from about 1.0 to about 15.0 weight percent of an metal oxide abrasive;
from about 0.3 to about 12.0 weight percent hydrogen peroxide or urea hydrogen peroxide;
from about 0.5 to about 3.0 weight percent ammonium oxalate; and benzotriazole.

20 34. The composition of claims 32 or 33, wherein the benzotriazole is present in an amount ranging from about 0.01 to about 0.2 weight percent.

35. The composition of claims 32-34, wherein the abrasive is alumina.

25 36. A chemical mechanical polishing slurry composition comprising:

an abrasive;
urea hydrogen peroxide;
from about 0.5 to about 3.0 weight percent ammonium oxalate; and
30 from about 0.01 to about 0.2 weight percent benzotriazole.

37. The composition of claim 36, wherein said composition includes from about 1.0 to about 15.0 weight percent of an alumina abrasive and from about 0.3 to about 12.0 weight percent urea hydrogen peroxide.

35 38. A chemical mechanical polishing slurry composition comprising:

from about 1.0 to about 15.0 weight percent of an alumina abrasive;
from about 0.3 to about 12.0 weight percent urea hydrogen peroxide;
from about 0.5 to about 3.0 weight percent tartaric acid; and
40 from about 0.01 to about 0.2 weight percent benzotriazole, the composition having a pH of from about 4.0 and 9.0.

39. The composition of claims 1-4, 6-7 or 9-38, which further includes at least one surfactant.

40. A method for polishing a substrate including at least one metal layer comprising the steps of:

45 (a) admixing the composition of claims 1-8 or 21-28 and deionized water to form a chemical mechanical polishing slurry precursor composition;
(b) admixing the precursor composition of step (a) with hydrogen peroxide to form a chemical mechanical polishing slurry composition;
50 (c) applying the chemical mechanical polishing slurry composition of step (b) to the substrate; and
(d) removing at least a portion of the metal layer from the substrate by bringing a pad into contact with the substrate and moving the pad in relation to the substrate.

41. The method of claim 40, wherein the hydrogen peroxide of step (b) is admixed with the precursor composition of step (a) at a ratio from about 0.75 moles of urea to 1 mole of hydrogen peroxide to about 2 moles of urea to about 1.0 moles of hydrogen peroxide.

42. A method for polishing a substrate comprising the steps of:

EP 0 846 742 A2

- (a) applying the chemical mechanical polishing slurry composition of claims 9-39 to the substrate; and
- (b) removing at least a portion of the metal layer from the substrate by bringing a pad into contact with the substrate and moving the pad in relation to the substrate.

5 43. The method of claims 40-42, wherein the substrate includes a copper alloy containing layer.

44. The method of claim 40-43, wherein the substrate further includes a titanium and titanium nitride layer wherein at least a portion of the titanium and titanium nitride layer is removed.

10 45. The method of claims 40-44, wherein the chemical mechanical polishing slurry composition is applied to the pad before the pad is placed into contact with the substrate.

46. A chemical mechanical polishing composition multi-package system comprising:

- 15
- (a) a first container comprising the composition of claims 1-8 or 21-28;
 - (b) a second container comprising hydrogen peroxide.

47. A chemical mechanical polishing composition multi-package system comprising:

- 20
- (a) a first container including a film forming agent and a complexing agent;
 - (b) a second container comprising an oxidizing agent; and
 - (c) an abrasive located in a container selected from the group consisting of the first container, the second container, or a third container.

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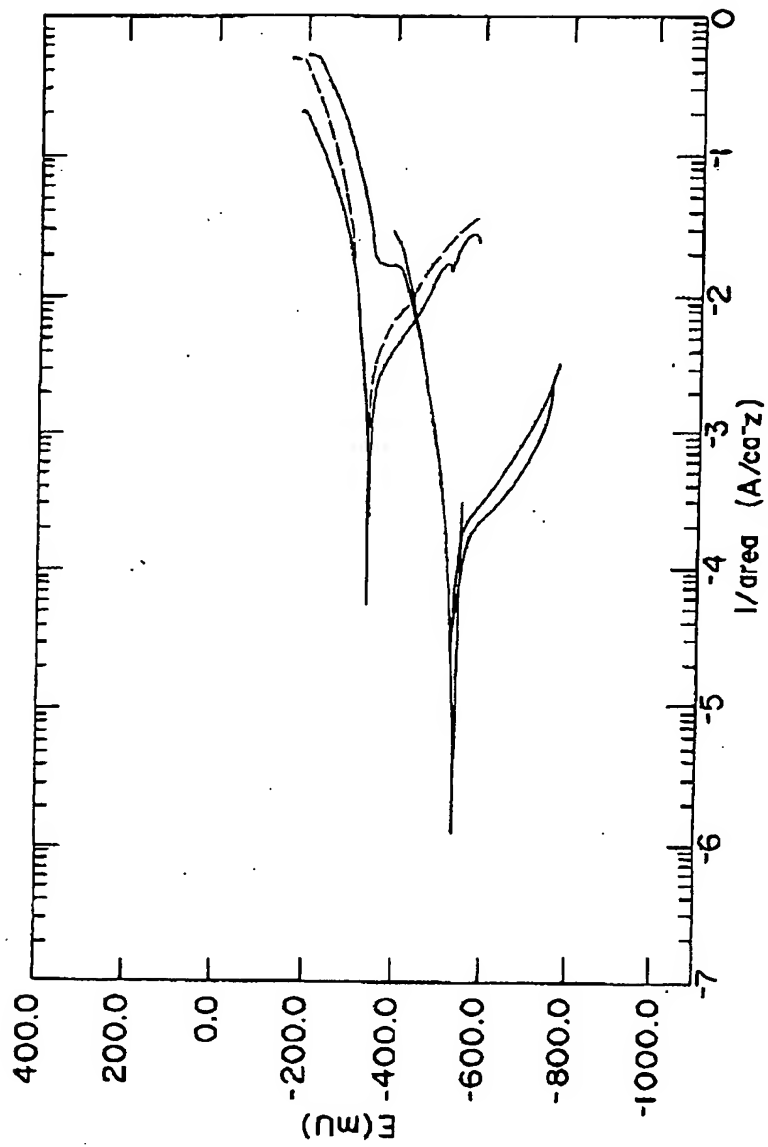


FIG. 1

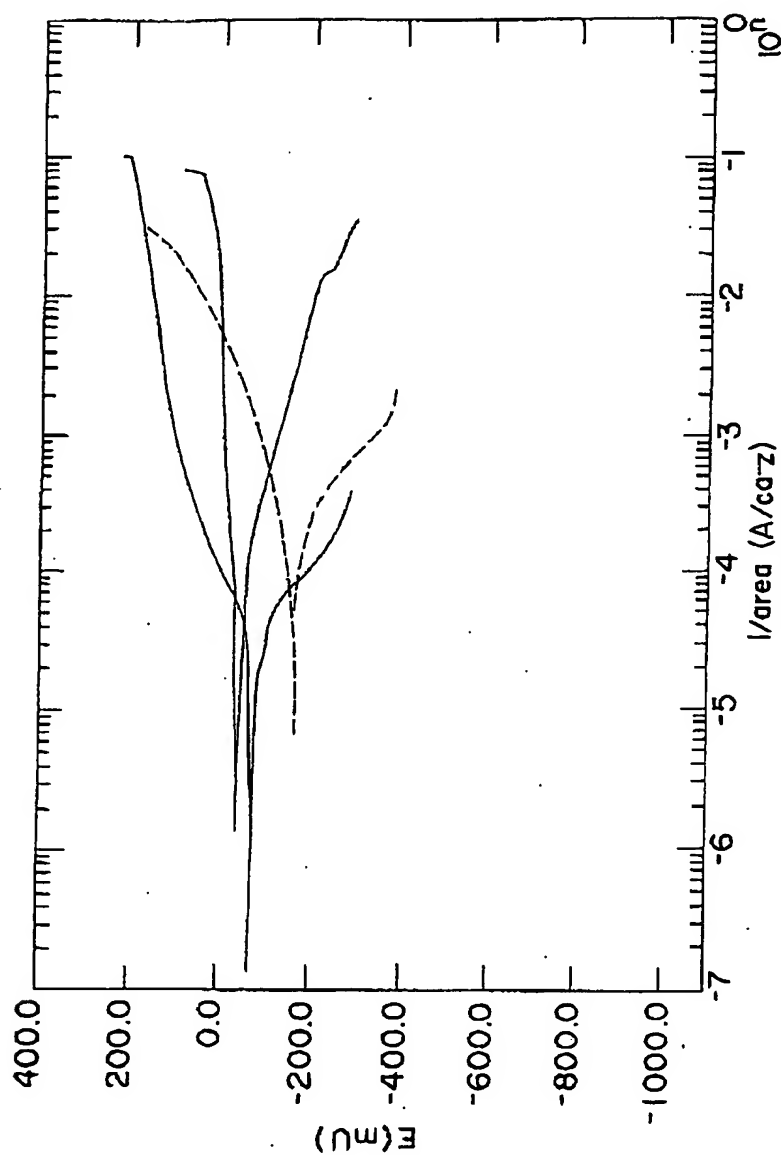


FIG. 2

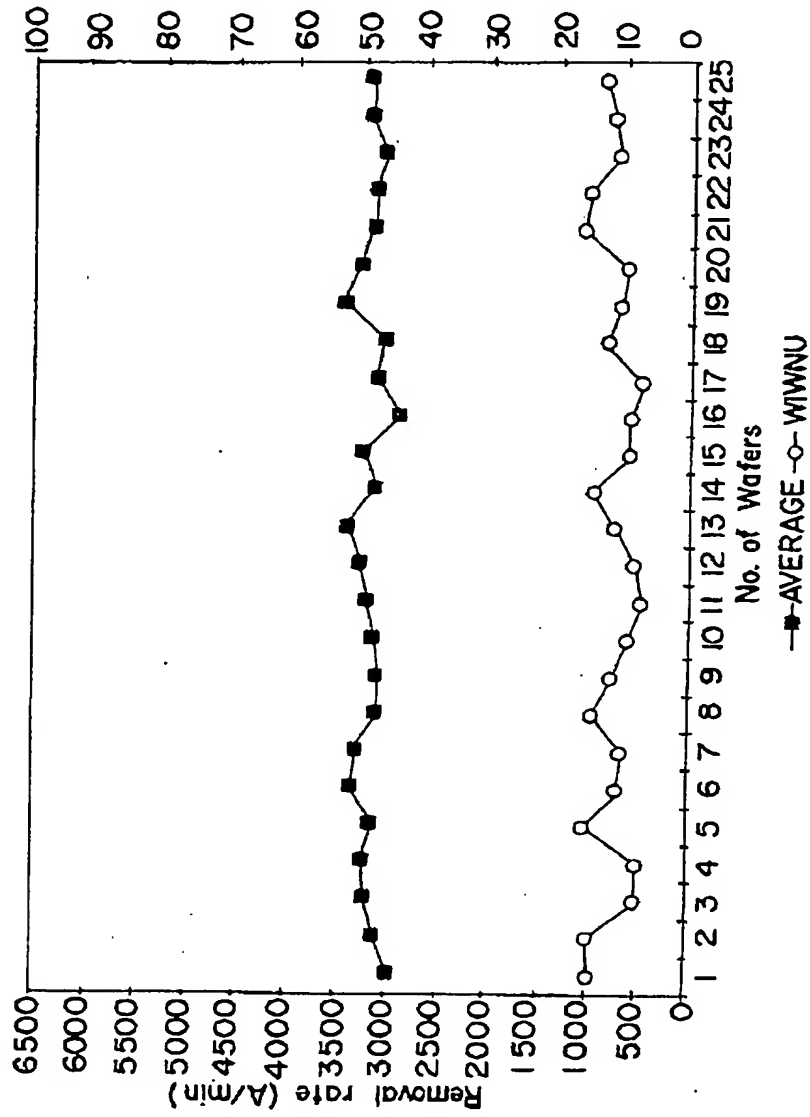


FIG. 3



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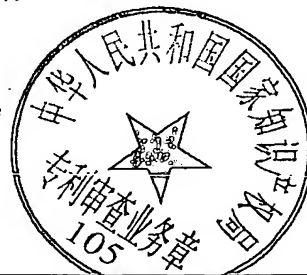
北京市朝阳区裕民路 12 号中国国际科技会展中心 A1210 号

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申请人:

日立化成工业株式会社

发明名称:

化学机械研磨用研磨剂及基板的研磨法



第一次审查意见通知书

(进入国家阶段的 PCT 申请)

1. ☒ 依申请人提出的实审请求, 根据专利法第 35 条第 1 款的规定, 审查员对上述发明专利申请进行实质审查。

☐ 根据专利法第 35 条第 2 款的规定, 国家知识产权局决定自行对上述发明专利申请进行审查。

2. ☒ 申请人要求以其在:

JP 专利局的申请日 1999 年 8 月 17 日 为优先权日,

专利局的申请日 为优先权日,

专利局的申请日 为优先权日,

3. ☐ 申请人于____年____月____日提交的修改文件, 不符合专利法实施细则第 51 条的规定。

☐ 申请人提交的下列修改文件不符合专利法第 33 条的规定:

☐ 国际初步审查报告附件的中文译文。

☐ 依据专利合作条约第 19 条规定所提交的修改文件的中文译文。

☐ 依据专利合作条约第 28 条或 41 条规定所提交的修改文件。

☐

4. ☐ 审查是针对原始提交的国际申请的中文译文进行的。

☒ 审查是针对下述申请文件进行的:

说明书 第____页, 按照原始提交的国际申请文件的中文译文;

第____页, 按照国际初步审查报告附件的中文译文;

第 1-12 页, 按照依据专利合作条约第 28 条或 41 条规定所提交的修改文件;

第____页, 按照依据专利法实施细则第 51 条规定所提交的修改文件。

权利要求 第____项, 按照原始提交的国际申请文件的中文译文;

第____项, 按照依据专利合作条约第 19 条规定所提交的修改文件的中文译文。

第____项, 按照国际初步审查报告附件的中文译文;

第 1-18 项, 按照依据专利合作条约第 28 条或 41 条所提交的修改文件;

第____项, 按照依据专利法实施细则第 51 条规定所提交的修改文件。

附图 第____页, 按照原始提出的国际申请文件的中文译文;

第____页, 按照国际初步审查报告附件的中文译文;

第 1 页, 按照依据专利合作条约第 28 条或 41 条所提交的修改文件;

第____页, 按照依据专利法实施细则第 51 条规定所提交的修改文件。





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5. ☒ 本通知书引用下述对比文献 (其编号在今后的审查过程中继续沿用):

编号	文件号或名称	公开日期 (或抵触申请的申请日)
1	EP0846742A2	1998年6月10日
2		
3		
4		

6. 审查的结论性意见:

☐ 关于说明书:

☐ 申请的内容属于专利法第5条规定的不授予专利权的范围。

☐ 说明书不符合专利法第26条第3款的规定。

☐ 说明书不符合专利法第33条的规定。

☐ 说明书的撰写不符合专利法实施细则第18条的规定。

☒ 关于权利要求书:

☐ 权利要求____不具备专利法第22条第2款规定的新颖性。

☒ 权利要求 2, 8-14 不具备专利法第22条第3款规定的创造性。

☐ 权利要求____不具备专利法第22条第4款规定的实用性。

☐ 权利要求____不符合专利法第26条第4款的规定。

☐ 权利要求____不符合专利法第31条第1款的规定。

☐ 权利要求____不符合专利法第33条的规定。

☐ 权利要求____不符合专利法实施细则第13条第1款的规定。

☐ 权利要求____不符合专利法实施细则第2条第1款关于发明的定义的规定。

☒ 权利要求 15-18 不符合专利法实施细则第20条至第23条的规定。

☐

上述结论性意见的具体分析见本通知书的正文部分。

7. 基于上述结论性意见, 审查员认为:

☐ 申请人应按照通知书正文部分提出的要求, 对申请文件进行修改。

☒ 申请人应在意见陈述书中论述其专利申请可以被授予专利权的理由, 并对通知书正文部分中指出的不符合规定之处进行修改, 否则将不能授予专利权。

☐ 专利申请中没有可以被授予专利权的实质性内容, 如果申请人没有陈述理由或者陈述理由不充分, 其申请将被驳回。

☐

8. 申请人应注意下述事项:

(1) 根据专利法第37条的规定, 申请人应在收到本通知书之日起的 肆 个月内陈述意见, 如果申请人无正当理由逾期不答复, 其申请将被视为撤回。

(2) 申请人对其申请的修改应符合专利法第33条的规定, 修改文本应一式两份, 其格式应符合审查指南的有关规定。

(3) 申请人的意见陈述书和/或修改文本应邮寄或递交给中国专利局受理处, 凡未邮寄或递交给受理处的文件不具备法律效力。

(4) 未经预约, 申请人和/或代理人不得前来中国专利局与审查员举行会晤。

9. 本通知书正文部分共有 2 页, 并附有下列附件:

☒ 引用的对比文件的复印件共 1 份 20 页。

☐

审查 9 部 5 室

审查员签章: 9583



完成日期: 2003-09-23



第一次审查意见通知书正文

如说明书所述，本申请涉及一种化学机械研磨用研磨剂及基板的研磨法。
经审查，现提出如下审查意见。

1. 权利要求 2 要求保护化学机械研磨用研磨剂，对比文件 1（说明书第 4 页 35 行-第 7 页 49 行）公开了一种用于铜基合金的化学机械抛光浆液，包含金属表面膜形成剂、金属氧化剂、酸、水和磨料，金属氧化的浓度在 1.0—12.0Wt%，化学机械研磨浆液的 pH 在 2.0-12.0 之间，具体的 pH 值可以通过掩膜浆液中酸、碱等来调节。由此可见，对比文件 1 已经公开了权利要求 1 的大部分技术特征，该权利要求所要求保护的技术方案与对比文件 1 所公开的技术方案相比，其区别在于：限定了化学机械研磨剂的 pH 值在 3 以下。通过调节研磨剂中的酸和碱的含量来调节其 pH 值，使得研磨剂溶液的 pH 值在 3 以下，是本领域技术人员所采用的常用手段，不具有突出的实质性特点和显著进步，因此权利要求 1 所要求保护的技术方案不具有专利法第 22 条第 3 款规定的创造性。
2. 权利要求 8-14 限定部分的附加技术特征同样被对比文件 1 所公开，对比文件 1（说明书第 4 页 35 行-第 7 页 49 行）公开了研磨剂中的酸为酒石酸；金属保护膜形成剂包括苯并三唑；氧化剂包括过氧化氢；导体是铜、铜合金；防止铜原子扩散的包括钽、钽合金；因此当其引用的权利要求没有创造性时，权利要求 8-14 所要求保护的技术方案不具有专利法第 22 条第 3 款规定的创造性。
3. 权利要求 15 要求保护一种化学机械用的研磨剂，该权利要求所要求保护的研磨剂，没有以研磨剂的化学成分，各个化学成分在研磨剂所占的比例来限定，而是以研磨剂的使用条件和性能来描述。另外该权利要求的技术特征出现了括号，这使得该权利要求的保护范围不清楚，不符合专利法实施细则第 20 条第 1 款的规定。
4. 权利要求 16 限定部分的附加技术特征中出现了括号，这使得该权利要求的保护范围不清楚，不符合专利法实施细则第 20 条第 1 款的规定，申请人应当将括号及括号的内容删除。
5. 权利要求 17、18 要求保护一种基板的研磨方法，方法权利要求应当工艺过、

操作条件、步骤或者流程等技术特征来描述。该权利要求所要求保护的技术方案没有记载所述基板研磨方法的工艺过、操作条件、步骤或者流程等技术特征，这使得该权利要求的保护范围不清楚，不符合专利法实施细则第 20 条第 1 款的规定。

6. 说明书第 2 页 4 行：“当这一过程反复进行时”应当为“当这一过程反复进行时”。

基于上述理由，本申请按照目前的文本是不能够被授权的。申请人应根据上述审查意见在指定的期限内提交新的权利要求书和/或说明书，修改时应满足专利法第三十三条的规定，不得超出原说明书和权利要求书记载的范围，如果申请人不能在本通知书规定的答复期限内克服上述缺陷或表明其具有符合所述规定的充分理由，本申请将被驳回。申请人应提供修改所涉及的原文复印件，并将修改之处用彩笔标示清楚。